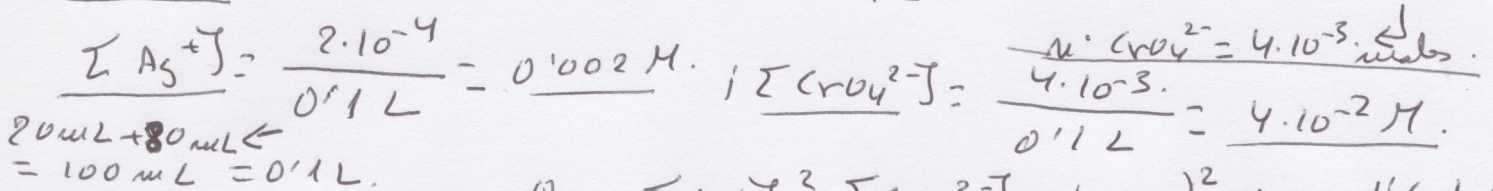
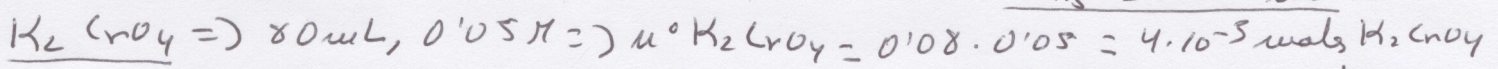
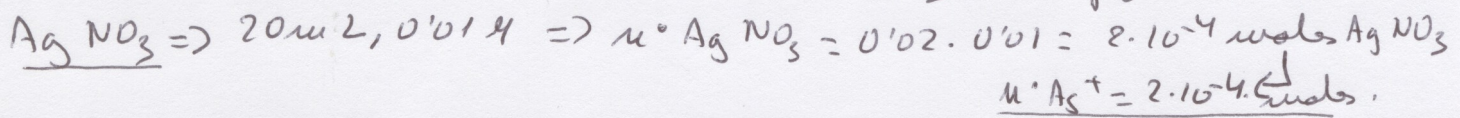
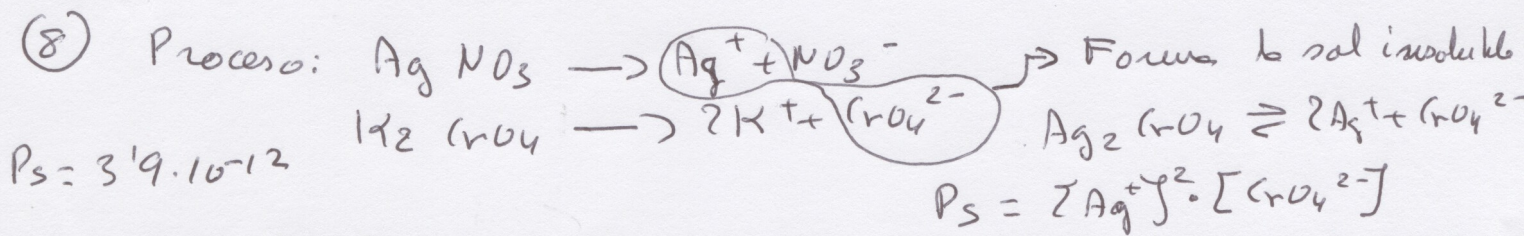
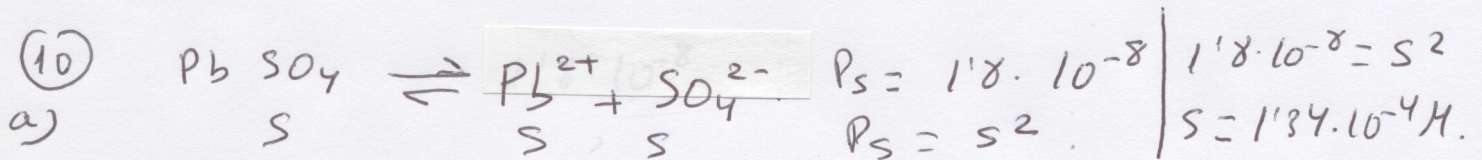


Q. 8. Producto de Solubilidad

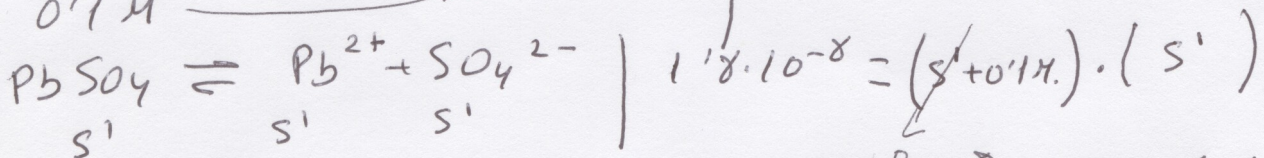
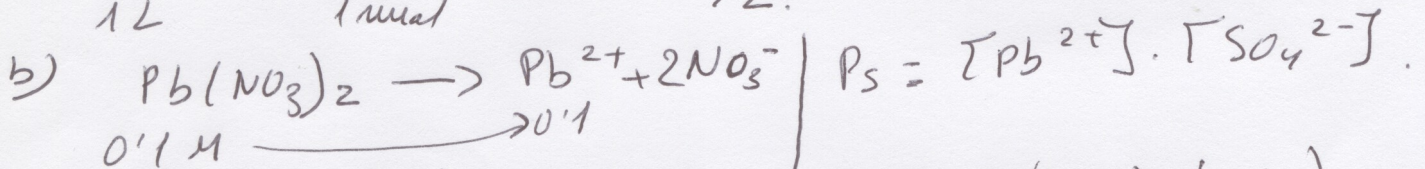


$$Q_s = [\text{Ag}^+]^2 \cdot [\text{CrO}_4^{2-}] = (0.002)^2 \cdot 0.04 = 1.6 \cdot 10^{-7}$$

$Q_s > P_s \Rightarrow$ Forma precipitado.

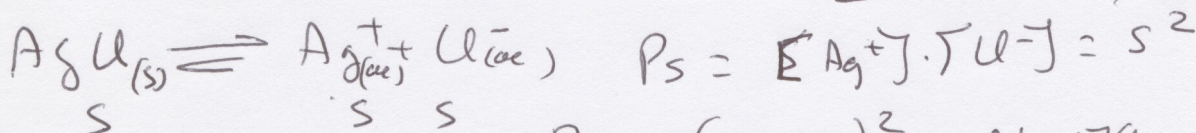
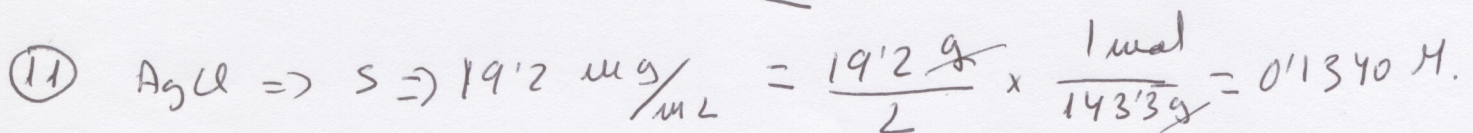


$$\frac{1.34 \cdot 10^{-4} \text{ mol}}{1 \text{ L}} \times \frac{303 \text{ g}}{1 \text{ mol}} = 0.041 \text{ g/L}$$



$$S' = \frac{1.8 \cdot 10^{-8}}{0.1} = 1.8 \cdot 10^{-7} \text{ M}$$

$$\frac{1.8 \cdot 10^{-7} \text{ mol}}{1 \text{ L}} \times \frac{303 \text{ g}}{\text{mol}} = 5.45 \cdot 10^{-5} \text{ g/L}$$



$$P_s = (0.1340)^2 = 0.0179$$

en disolución de CaCl_2 0.8 M

